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| **Instrument Number 1** | **Term 3 2022** |
| **Student Name** |  | **Handout Date** (Week Beginning) | 22/08/2022 |
| **Teacher Name** |  GILLMU, TURNGA | **Interim Check Date** | Ongoing |
| **Unit Number/Name** | Chemistry Preparation | **Rough Draft Date** | Ongoing |
| **Due Date** | 12/09/2022 |

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| **Assessment Technique** | Student Experiment |
| **Time/Length** | 3 weeks |
| **Assessment Conditions** | Summative |
| **Seen/Unseen** | Seen and unseen elements |
| Materials handed out prior to assessment? | No [ ]  | Yes [x]  | **Conditions** |
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| **Criterion** | **Marks** |  |
| **Science Understanding** | /20 |  |

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| **Differentiation: If assessment conditions have been adjusted details are provided below** |
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| **Acknowledgement of assessment responsibility** |  |
| I understand the consequences of plagiarism/cheating and confirm this is my own work. |
| **Student Signature:**  | **Date:** ……………………………… |

**Criterion: Communication**

Assessment Objective

7. communicate understandings and experimental findings, arguments and conclusions about reaction rates and collision theory

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| **The student work has the following characteristics:** | **Marks** |
| • effective communication of understandings, findings, arguments and conclusions about reaction rates and collision theory demonstrated by­ fluent and concise use of scientific language and representations­ appropriate use of genre conventions­ acknowledgement of sources of information through appropriate use of referencing conventions. | 2 |
| • adequate communication of understandings, findings, arguments and conclusions about reaction rates and collision theory demonstrated by­ competent use of scientific language and representations­ use of basic genre conventions­ use of basic referencing conventions. | 1 |
| • does not satisfy any of the descriptors above. | 0 |

**Criterion: Research and Planning**

Assessment Objectives

2. apply understanding of reaction rates and collision theory to modify experimental methodologies and process primary data

5. investigate factors that effect rates of reaction through an experiment

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| **The student work has the following characteristics** | **Marks** |
| • informed application of understanding of reaction rates and collision theory to modify experimental methodologies demonstrated by:­ a considered rationale for the experiment­ justified modifications to the methodology • effective and efficient investigation of rates of reaction demonstrated by:­ a specific and relevant research question­ a considered methodology that enables the collection of sufficient, relevant data­ considered management of risks and ethical or environmental issues. | 5–6 |
| • adequate application of understanding of reaction rates and collision theory to modify experimental methodologies demonstrated by­ a reasonable rationale for the experiment­ feasible modifications to the methodology• effective investigation of reaction rates and collision theory demonstrated by­ a relevant research question­ a methodology that enables the collection of relevant data­ management of risks and ethical or environmental issues. | 3–4 |
| • rudimentary application of reaction rates and collision theory demonstrated by­ a vague or irrelevant rationale for the experiment­ inappropriate modifications to the methodology• ineffective investigation of reaction rates and collision theory demonstrated by­ an inappropriate research question­ a methodology that causes the collection of insufficient and irrelevant data­ inadequate management of risks and ethical or environmental issues. | 1–2 |
| • does not satisfy any of the descriptors above. | 0 |

**Criterion: Analysis of Evidence**

Assessment Objectives

2. apply understanding of reaction rates and collision theory to modify experimental methodologies and process primary data

3. analyse experimental evidence about reaction rates and collision theory

5. investigate reaction rates and collision theory through an experiment

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| **The student work has the following characteristics:** | **Marks** |
| • appropriate application of algorithms, visual and graphical representations of data about reaction rates and collision theory demonstrated by­ correct and relevant processing of data• systematic and effective analysis of experimental evidence about of reaction rates and collision theory demonstrated by­ thorough identification of relevant trends, patterns or relationships­ thorough and appropriate identification of the uncertainty and limitations of the evidence• effective and efficient investigation of reaction rates and collision theory demonstrated by ­the collection of sufficient and relevant raw data. | 5–6 |
| • adequate application of algorithms, visual and graphical representations of data about reaction rates and collision theory demonstrated by­ basic processing of data• effective analysis of experimental evidence about reaction rates and collision theory demonstrated by­ identification of obvious trends, patterns or relationships­ basic identification of uncertainty and limitations of evidence• effective investigation of reaction rates and collision theory demonstrated by ­ the collection of relevant raw data. | 3–4 |
| • rudimentary application of algorithms, visual and graphical representations of data about reaction rates and collision theory demonstrated by ­ incorrect or irrelevant processing of data• ineffective analysis of evidence demonstrated by­ identification of incorrect or irrelevant trends, patterns or relationships­ incorrect or insufficient identification of uncertainty and limitations of evidence• ineffective investigation of reaction rates and collision theory demonstrated by ­ the collection of insufficient and irrelevant raw data. | 1–2 |
| • does not satisfy any of the descriptors above. | 0 |

**Criterion: Interpretation and Evaluation**

Assessment Objectives

 4. interpret experimental evidence about reaction rates and collision theory

 6. evaluate experimental processes and conclusions about reaction rates and collision theory

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| **The student work has the following characteristics:** | **Marks** |
| • insightful interpretation of experimental evidence about reaction rates and collision theory demonstrated by­ justified conclusion/s linked to the research question• critical evaluation of experimental processes about reaction rates and collision theory demonstrated by­ justified discussion of the reliability and validity of the experimental process­ suggested improvements and extensions to the experiment which are logically derived from the analysis of the evidence. | 5–6 |
| • adequate interpretation of experimental evidence about reaction rates and collision theory demonstrated by­ reasonable conclusion/s relevant to the research question• basic evaluation of experimental processes about reaction rates and collision theory demonstrated by­ reasonable description of the reliability and validity of the experimental process­ suggested improvements and extensions to the experiment which are related to the analysis of the evidence. | 3–4 |
| • invalid interpretation of experimental evidence about reaction rates and collision theory demonstrated by ­ inappropriate or irrelevant conclusion/s• superficial evaluation of experimental processes about reaction rates and collision theory demonstrated by­ cursory or simplistic statements about the reliability and validity of the experimental process­ ineffective or irrelevant suggestions. | 1–2 |
| • does not satisfy any of the descriptors above. | 0 |

**ORIGINAL EXPERIMENT**

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| **Reaction Rate between Marble Chips and Hydrochloric Acid measured through mass loss** |

Marble chips (calcium carbonate) react with dilute hydrochloric acid according to the equation:

CaCO3 (s) + 2 HCl (aq) ⎯→ CaCl2 (aq) + CO2 (g) + H2O (l)

While the reaction is proceeding, the mass of the reaction flask (with its contents) decreases as gas escapes.

By observing the loss in mass (which is the mass of carbon dioxide gas evolved), we can investigate how the reaction rate changes during the course of the reaction.

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| **Materials** | **Chemicals** |
| * 100 mL conical flask
 | 50 mL of 2.0 mol L-1 Hydrochloric acid |
| * Cotton wool to be used as a plug
 | 2 grams Calcium carbonate chips |
| * Electronic balance
 |  |
| * Filter paper
 |  |
| * 100 mL graduated cylinder
 |  |
| * 100 mL beaker
 |  |
| * Stopwatch
 |  |

**For each of the apparatuses determine the measurement uncertainty associated with the instrument**

100 mL graduated cylinder ± \_\_\_\_\_\_\_\_\_\_ mL

Electronic balance ± \_\_\_\_\_\_\_\_\_\_ g

Stopwatch ± \_\_\_\_\_\_\_\_\_\_ s

**METHOD**

1. Use a 100 mL graduated cylinder measure out 50 mL of 2.0 mol L-1 hydrochloric acid into a 100 mL conical flask.

**Figure 1:** Diagram of experiment set up

1. Record the exact volume of hydrochloric acid used \_\_\_\_\_\_\_\_\_\_\_\_\_\_ mL
2. Plug the neck of the flask loosely with cotton wool.
3. Weigh out 2.00 g of large marble chips on a filter paper on electronic balance.
4. Record the exact weight of calcium carbonate \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ g.
5. Put the flask on the pan of the balance as seen in Figure 1.
6. Record the **total mass** and record the reading in the table 1 in the results section as time zero (0).
7. Remove the cotton wool plug. Add the marble chips to the acid and quickly replace the plug. Start the stopwatch immediately. (Leave the flask with its contents and the filter paper on the balance pan.)
8. Record the total mass every 10 seconds for the first two minutes and every half-minute for the rest of the time, until two **or three consecutive constant readings have been obtained**
9. Calculate the total loss in mass for each time **Total loss in mass = mass recorded – initial mass**

**RESULTS**

Table 1: Mass of experiment vs time (seconds)

|  |  |  |  |  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- |
| Time (s) ±…………. *(uncertainty)* | 0 | 10 | 20 | 30 | 40 | 50 | 60  | 70 | 80 | 90 | 100 |
| Total mass (g)±…………. *(uncertainty)* |  |  |  |  |  |  |  |  |  |  |  |
| Total loss in mass (g)±…………. *(uncertainty propagated)* | 0 |  |  |  |  |  |  |  |  |  |  |
| Time (s) ±…………. *(uncertainty)* | 110 | 120 | **150** | **180** | **210** | **240** | **270** | **300** | **330** | **360** | **390** |
| Total mass (g)±…………. *(uncertainty)* |  |  |  |  |  |  |  |  |  |  |  |
| Total loss in mass (g)±…………. *(uncertainty propagated)* |  |  |  |  |  |  |  |  |  |  |  |

1. Plot a graph of total loss in mass (g) against time (seconds). Draw a smooth curve through as many points as possible.

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**Answer the following questions:**

1) Identify the time period at which the rate of reaction is the fastest.\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

2) Identify the time period at which the rate of reaction is the slowest. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

3) Determine the total mass loss of carbon dioxide \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

4) Calculate the moles of calcium carbonate you used in the experiment.

5) Calculate the moles of hydrochloric acid you used in the experiment.

6) Which reagent (calcium carbonate, or Hydrochloric acid) is the limiting reagent? Show your workings

7) Calculate the theoretical yield of carbon dioxide gas.

CaCO3 (s) + 2 HCl (aq) ⎯→ CaCl2 (aq) + CO2 (g) + H2O (1)

8) Determine the average **rate of the reaction =**$ \frac{Δm}{Δ t}$ by following the outline shown below in the table.

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| STEP 1: Calculate the change in mass and its uncertainty for a selected time period. Time period chosen was 0 sec to \_\_\_\_\_ sec. |
| Calculation | Measurement Uncertainty (primary data) | Random Uncertainty (primary data) |
| Change in mass = mf - mi = -  = g | Change in mass was calculated by subtracting, so use the uncertainty rule for propagating uncertainty for subtraction.That is, add the measurement uncertainty associated with both mass values.= ± \_\_\_\_\_ g + ± \_\_\_\_\_ g= ± \_\_\_\_\_ g  | Cannot do this as you only did one trial. Can only calculate random uncertainty when there are multiple trials. You will be able to do this when you design your own experiment as you will include multiple trials in your experiment.$$Random uncertainty = \frac{range}{2}$$ |
|  | If you had calculated both the measurement uncertainty AND the random uncertainty – you would now identify which of these uncertainties is greatest, and this amount of uncertainty is used is all further steps. |
| Change in mass = ± g |

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| STEP 2: Calculate the change in time  |
| Calculation | Measurement Uncertainty (primary data) | Random Uncertainty (primary data) |
| Change in time*Δt* = tf - ti= - = s | Change in time was also calculated by subtracting, so use the uncertainty rule for propagating uncertainty for subtractionThat is add the measurement uncertainty associated with both time values.= ± \_\_\_\_\_ s + ± \_\_\_\_\_ s= ± \_\_\_\_\_ s  | Cannot do this as you only did one trial. Can only calculate random uncertainty when there are multiple trials. You will do this when you design your own experiment.$$Random uncertainty = \frac{range}{2}$$ |
|  | If you had calculated both the measurement uncertainty AND the random uncertainty – you would now identify which of these uncertainties is greatest, and this amount of uncertainty is used is all further steps. |
| Change in time = ± s |
| STEP 3: Calculate the rate of the reaction  |
| Calculation | Uncertainty (secondary data) |
| Calculate the average rate of reaction$$ Rate= \frac{change in mass}{time}$$ = $ \frac{ g}{ s}$ = g/s | Rate of reaction (see left) was calculated by dividing, so use the uncertainty rule for propagating uncertainty for division. Note that as we are generating secondary data, there is no longer a measurement or random uncertainty, just uncertainty. This is calculated using the highest uncertainty values you have for the primary data. There is a formula you can use, or a series of steps. The formula is the process used below. $$Uncertainty =\pm \frac{change in mass}{time} (\frac{ uncertainty in change in mass}{change in mass}+ \frac{ uncertainty in time}{time}) $$$$Uncertainty =\pm g/s$$ |
| **Rate of reaction =** $\pm $ **g s-1** |

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|  **Task** |
| * Modify (i.e. refine, extend or redirect) the given investigation: “The rate of reaction between marble chips and hydrochloric acid **”.**
 |
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| **Scaffolding**  |
| The response must be presented using an appropriate scientific genre of a scientific report and contain the following parts (use appropriate heading):* **Rationale**
* **Research question**
* **Original Experiment and Modifications to the methodology**
* **Risk management and Environmental Issues**
* **Raw Data**
* **Data Processing**
* **Trends, Patterns and Relationships**
* **Limitations and Improvements**
* **Evaluation of the RELIABILITY and VALIDITY of the experimental process**
* **Conclusion**
* **Reference list**
* **Word Count**

**Rationale**: The rationale is where you give meaning to why you are performing the experiment. It should briefly give a real world context relating to your Research question. Any important theory concepts that relate to your investigation should be mentioned here. |

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| **Research question** - Developing a researchquestion |
| State the research question in one sentence.The independent and dependent variable should be clearly stated in the research question. The range of variation in the independent variable should also be stated, along with any particular context the research question is based in.EXAMPLE - How does changing *(Independent Variable – with description of range )* in the*(experiment or context)* affect the *(dependent Variable)* ?  |
| **Original Experiment**Brief reference to original experiment - authors of the study/experiment, the intent of the original investigation, and main findings.  |
| **Modifications to the methodology*** Explain (justify) how the experiment you are conducting is a modification of the original methodology in terms of how it will extend and/or refine the findings of original experiment.
* Describe each modification and justify why it is occurring.
* Clearly explain how you will be collecting sufficient, relevant data. Recommended there be a minimum of 5 variations to the independent variable determine a trend and a minimum of 3 trialsfor each independent variable.

Extended means:* Development of the experiment to further investigate …

Refined means:* Use of more precise measurement or instruments in order to get a more accurate answer, be specific.
 |
| **Risk Management and Environmental Issues**Careful thought has gone into identifying the potential hazards/risks for each chemical used and strategies to handle these risks. (use the following table as a guide give 3 potential hazards/risks and strategies to handle these)Risk Management:

|  |  |  |  |
| --- | --- | --- | --- |
| Source of risk | Amount of harm could it cause? (circle) | Safety precautions taken | If an incident occurred what should I do? |
|  | Minor Significant major |  |  |
|  | MinorSignificantmajor |  |  |
|  | MinorSignificantmajor |  |  |

Environmental Issues:Discuss any environmental issues that arise from using or making chemicals in your experiment. |

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| **Raw data -** Quantitative (numerical) data should be in Tables; qualitative (descriptive) data can be written in sentencesExample:Table 1: Description of data ( Design your own table **)**

|  |  |  |
| --- | --- | --- |
| **Independent variable** (units) | Measurement uncertainty | **Dependent variable** (units) |
| **Trial 1**  | **Trial 2** | **Trial 3** | **Average** **Δ mass** | Measurement uncertainty | Random Uncertainty |
| **Initial** | **final** | **change** | **Initial** | **final** | **change** | **Initial** | **final** | **change** |
| **2.0** |  |  |  |  |  |  |  |  |  |  |  |  |  |
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**Data Processing** – showing an example of each step in the calculation of your secondary data, alongside the propagation of error for each step. Compared to the first experiment there is an extra step at the start - for the independent variable, rest are for the dependent variable.Example:

|  |
| --- |
| STEP 1: Calculate the ……………………………………….. (using data set ………………………..) |
| Calculation | Measurement Uncertainty (primary data) | Random Uncertainty (primary data) |
|  |  |  |
| NOTE: *Highest uncertainty value is ±*  ………………..………………………..…**= ±**  |
| STEP 2: Calculate the …………………………………… (using data set ………………………..) |
| Calculation | Measurement Uncertainty (primary data) | Random Uncertainty (primary data) |
|  |  |  |
| NOTE: *Highest uncertainty value is ±*  ………………..………………………..…**= ±**  |
| STEP 3: Calculate the ……………………………….  |
| Calculation | Uncertainty (secondary data) |
|  |  |
| **…………………………….. =** $\pm $g s-**1** |

**Secondary Data:** Present your final data in a table. Table \_\_ : *Description of data*

|  |  |  |
| --- | --- | --- |
| **Independent Variable** (units) | **Dependant variable** (units) | **Uncertainty (%)** |
|  |  |  |
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Graphical analysis of dataSecondary (processed) data - Graph dependent variable (average rate) versus “independent variable” (will vary depending on what you choose for your independent variable) (Graph should be a scatterplot, with an appropriate scale and a trend line, with equation and R2 value displayed)**Identifying Trends, patterns and relationships**Generally speaking trends and relationships are the same idea (very unusual to “have patterns” in Chem). Both words can be used to identify or describe how the dependent variable (on the y axis) changes when independent variable (on the x axis) changes. However, the word relationship is more useful when talking about the maths that connects two variables. For this section write a paragraph that:* Starts with a very simple sentence identifying the trend in the data. The sentence should go along the lines of…

“As the (*independent variable*) increases, the (*dependent variable*) increases/decreases (*use which ever applies*)”.* The second part of the paragraph is a description of the trend. This can be simple, or more complex, depending on the data (shape of the trend line), and how much experience you have. Describe the trend as linear, exponential, polynomial, power, or logarithmic.
* The 3rd part of your paragraph will be used to quote data from the table or graph which supports the description of the trend.
* At the end of the paragraph you can explain any specific implication of this trend. Your trends may not have a specific implication, especially if they are simple trends like linear trends.

**Identifying Uncertainty and Limitations** Do this is two parts – Uncertainty, and LimitationsUncertainty describes how certain you are that your data is precise. * Start with a simple statement stating the amount of uncertainty in the data.

“There is a (*use an appropriate word*) amount of uncertainty in the data”. The rest of the paragraph is simply justifying why you think there is this much uncertainty. Can do this three ways. Two ways look at uncertainty in the data, one at the uncertainty in the trend. You need to apply all three ways. There is uncertainty in the data if only one of these ways identifies it.* The first way to identify uncertainty is from the data values. Generally speaking, an uncertainty which is higher than five or ten percent (do the percent calculation if it is not shown) is considered to be significantly uncertain.
* The second way to identify uncertainty is to see if there was data which was not used in the final results. Scientists sometimes label them anomalies. These anomalies are not used to calculate averages. Because the anomalies have been removed, the final data set can seem to have low uncertainty. However, the fact that there are anomalies in the raw data means there is more uncertainty in the data.
* The third way to identify uncertainty is by looking at the data and the trendline on a graph. Data points which are very close to a trend line suggests that the trend is very precise description of the data. This means the trend has low uncertainty. If the data points are not close to the trend line, the trend is not a good description of the data, and there is a significant amount of uncertainty in the trend.

Limitations are things which mean your experiment process did not necessarily measure what it was intended to. Almost all data contains limitations, you just have to identify them and explain what they are. * Start this section with a simple statement along the lines of

“ There appear to be (*use an appropriate word*) limitations in the data”. For the rest of the paragraph, you need to justify what these limitations are by identifying and briefly explaining each limitation. There are three ways to identify the limitations. You need to apply all three ways. * The first way to identify uncertainty is to look at how much data was measured. Generally speaking there needs to be 5 variations of the independent variable, and three trials for each variation. Five variations of the independent variable means there will be five data points on a graph. If there is less than five data points on the graph the data is limited in its ability to identify a trend. If there are no trials and no averages used, the data is also limited in its precision.
* The second way to identify limitations is to look at the value of the independent variable. Sometimes this is not a real-life value, but one the scientist used so the experiment/investigation could be completed in a shorter amount of time. Say a scientist was measuring the rate of decay in human skeletal remains – which may take many years to do in real life conditions. Scientists will often use an independent variable which is not real-life (for example - tiny bone fragments rather than whole bones) to speed things up. This is not “wrong”, and perfectly justified in science, but it can limit how well the data applies to real life conditions.
* The third way to identify limitations is by identifying any controlled variables which were not actually controlled. If controlled variables were not well controlled, this limits the accuracy of any trend which has been identified.

**Conclusion**The conclusion section is usually relatively short. The first paragraph has to be the main conclusion made from the data.* Start the paragraph with a simple statement of the actual conclusion you have made.
* Follow this statement with a description of the trend supporting the conclusion. If you have a mathematical equation, include it here to support your simple statement. If you do not have a mathematical equation, elaborate and further explain your conclusion or the trend which supports it.
* Explain the real-life implications of the conclusion. That is, explain how this conclusion is relevant and important. Doing this generally shows you understand the significance of the conclusion.

**Evaluation of the RELIABILITY and VALIDITY of the experimental process**This sounds complicated, but is actually quite easy. It is easier if you discuss reliability separately from validity, so do this in two parts.Reliability generally refers to whether your data is repeatable. That is, if anyone else was to do the investigation using the same method, would they get the same result? The uncertainty you identified earlier can be used here. If you have identified a lot of uncertainty, the method is not reliable. If you were unable to identify any uncertainty, then the data is reliable. Reliability should only take one paragraph, and:* Start the paragraph by stating if the method and resultant data is reliable, not reliable, or somewhere in between. Then justify this by explaining the uncertainty you identified earlier. Then try to identify why that uncertainty occurred by evaluating the errors made in the method. This last bit is the hard part.

 Validity generally refers to whether the trends you identified, and the conclusions you made are valid or not. That is: are the trends real, and are the conclusions meaningful? The limitations you identified earlier can be used to argue validity, but there is also one additional way to evaluate validity. This is by comparing your conclusions to existing theory and determining if there is error in your conclusion – traditional error analysis. Evaluating validity should also only take one paragraph, and the sequence you should follow is outlined below.* Start the paragraph by stating if the conclusions made in the investigation are valid, not valid, or somewhere in between. Use the limitations you identified earlier to justify your decision in the first statement. Explain how the limitations could have been avoided.

Word Count: 1000 to 1500 words**Reference List**Use Harvard referencing as prescribed in your student learning journal. |